War Record of Chas. C. Colby. Constant Prof of Geograf Special Expert, Commodity Section, Division of Planning and Stastics United States Shipping Board, 1918-19. Major work was on minor food commodities, Gentral and South American problems and Pacific Ocean tonnage. War Record of Chas. C. Colby. Conduct Prof of Quegra

Special Expert, Commodity Section, Division of Planning and Stastics United States Shipping Board, 1918-19. Major work was on minor food commodities, Gentral and South American problems and Pacific Ocean tonnage.

Name in full		John	Morle	Coulter		Date	anuary	16-1920
Academic title_		Profes	our and H	ead of the Department of Botany		U		
		DATE		(State in full, without abbreviation) RANK, BRANCH of SERVICE, and ORGANIZATION, specifying Co., Reg., and Div., or similar designation of unit	g	LOCATION	Abbi RANK	reviations for BRANCH
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The blank does not fit my particular case. My war connection was as Chairman of the Committee on Botany of the National Research Council. This committee worked on emergency problems connected with food production, food values, timbers, and raw materials (used in manufactures, in drugs, and for many other purposes).

Enlistment Promotions, 4 All Cardina in the without abbreviation of unit and the control of th	200 January 16 - 1920		Name in fullAcademic title
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Name in full William John Crogies	Date 28 th Jan. 20
Name in full William John Crogier Academic title Assistant Professor of Zoölogy.	

		DATE	(State in full, without abbreviation) RANK, BRANCH of SERVICE, and ORGANIZATION, specifying Co., Reg., and Div., or similar designation of unit		LOCATION	Abbi RANK	reviations for BRANCH
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[During Jan-Feb-March of 1915 I was research assistant at the andulause americane, Therilly, Pavis, France. From 15 to '18 I lived at the Bermuda Felands, and was not involved in military operations.

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Date 28 24 2m. 20				liam		Name In full.
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State in detail military exploits performed, honors and marks of distinction conferred (including service chevrons), and any items of interest concerning your service, such as position of instructor in special unbject, etc.

During

an-Jeb-March of 1915 I was herearch assistant at the andurence arrangement, Travilly, Paris France From 15 to 10 I lived at the Bermuda Folands, and was not involved in military operations

Name in full_	0	athu	Jeffrey Dempster	Date	Jan 1	7,1920
Academic title_			Instructor in Physics-			
		DATE	(State in full, without abbreviation) RANK, BRANCH of SERVICE, and ORGANIZATION, specifying Co., Reg., and Div., or similar designation of unit	LOCATION	Abbr	reviations for BRANCH
Enlistment	n	10523,1917	Private, Lignal Corps, attached 813th Depot are Rquadron Boston, assigned to	U.S. Naval	Experiment	Sig Gpr tal Station, New London, Con
	I M	104. 198	Romotes, Master Signal Electrician. 11 " "	New London New york Cty	M.S.E.	Sig Cos.
	3 1	ept 12, 1918	Promoted, Master Signal Electrician. " Transferred to 814th Depot aus Aquadron, New York City, assigned to Workern Electric Engineer Exometed Commissioned 2nd Lieut Dignal Corps transferred to Camp alfred Vail, L	attle Silver N.	Instigued J. assigned	to Officers Training Bathele
Promotions, Transfers, Offices held, e.g., Co. Comdr., Adj., etc.	4 00	723,19/8	Sailed for France. Attached to Research Division of signal Corps	Paris		Sig Eps.
	6 2	ec 25, 1918	attached to Scientific attacht, america Embassy, Ordered to Rome of Javanto, Italy			
	8 4	april 2, 1919	on detached herrice with the Italian Nay, later at Toulow with the French Way assigned as 3nd Lient to Casual Co 3481. anied in New York may 8, 1919.	-		
Discharge	4	nay 1919	Discharged at Camp mitchell Field, Lio?			
State in detail	military	ex-	non Mor 30 - June 22, 1918 I was engaged, in the development of and the	stong of Su	Imarine	detecting devices,

and later in the development and testing of methods of secret signalling. From June 22- Sept 12, 1918, I was working with wireless telephone apparatuse for aeroplanes.

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	LOCATION		
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	d Listers of All		

Prof. of Que. Heist.

History Department,

University of Chicago.

War Service Information:

I am not sure that I did anything that falls within the purview of war service office; but I am risking the following statement:

1, I have been spending my summers on a small farm near Harpers Ferry, Virginia, since 1913. There was a tenant on the farm when we entered the war. In order to increase the food output in 1918, I rented a piece of waste land adjoining and planted it in corn. We made 912 bushels The next year we planted wheat on part of it and 390 "The remainder was planted in corn or grazed, result ... 360 "

Total corn 1,272 bushels for the two years; wheat 390 bushels.

The extra cost to me was, rentals \$300 a year or \$600 total. Other expenses, estimated, \$200

2,0f less importance: In the autumn of 1917 and early winter of 1918 I prepared a preliminary study of the Japanese-Chinese situation in the far East and a similar study of the Monroe Doctrine at the request of Colonel House. These were duly submitted and accepted; but I do not know whether my conclusions were of any practical value in the scramble at Paris. Still a third survey was made for Colonel House personally. It was a study of recent American industrial and commercial tendencies. I mention these matters merely. They have no place in the kind of list your blank called for.

Yours sincerely,

William E. Dodd

William & nord

0			te Medial Cannon; Professional Lecturer (State in full, without abbreviation)		Abbr	eviations for
	DATE		RANK, BRANCH of SERVICE, and ORGANIZATION, specifying Co., Reg., and Div., or similar designation of unit	LOCATION	RANK	BRANCH
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The University of Chicago The School of Education DEPARTMENT OF NATURAL SCIENCE DEPARTMENT OF ZOÖLOGY Jason Prof. of Westernal Seems February 17, 1920. Mr. D. A. Robertson President's Office University of Chicago, Chicago, Illinois.

Dear Sir:

I take it that on the enclosed blank it is desired to have entered only actual military service. I am sorry that my age prevented my enlistment.

I was a member of the Illinois State Guard, was Director of the hygiene work of the Junior Red Cross in the Middle West, and did Four Minute Man speaking twice a week.

Yours very truly...

ERD.AF

Daiversity of Chicago,

Name in full	Cars	on Davis	ul Duncan	Date	2, 1920
	former	ey Assist	ant Professor of Commercial		
	DATE	RAN	(State in full, without abbreviation) K, BRANCH of SERVICE, and ORGANIZATION, specifying Co., Reg., and Div., or similar designation of unit	LOCATION	Abbreviations for RANK BRANCH
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	3 Sept. K		istician, American Shipping Misso	on, London	2. England
Promotions, Transfers, Offices held,	4 Feb.	27,1919.	Statistician, U.S. Shipping Book, T	Paris, Peace	
e.g., Co. Comdr., Adj., etc.	5	cons	levence_		
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	7 60	for runs &	mes, forly 15, 1919.		
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Name in full	Tinch, Edwin Freder	ick	Date_	amar	76,192
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	DATE RANK, BRANCH	State in full, without abbreviation) Tof SERVICE, and ORGANIZATION, specifying Total and Div., or similar designation of unit	LOCATION	Abbre RANK	eviations for BRANCH
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Discharge	april 29, 1919.	Base Stoopetul	Camp Trans	Caph	m.c.

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Seo Finkelsten Name in full Academic title (State in full, without abbreviation) Abbreviations for DATE RANK, BRANCH of SERVICE, and ORGANIZATION, specifying LOCATION RANK BRANCH Co., Reg., and Div., or similar designation of unit Private Enlistment 10 1) 11 1) Drance Jamo 11 11 Promotions, Transfers, Offices held, e.g., Co. Comdr., Adj., etc. Paris 11 11 10 11 11 33 .. 8

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- General Alt C. W. 8.	Discharge July 1919 2nd St. chaming Warpon Series

Name in full	- Fro	igr Enfield Frager	Date	Jan. 22, 1920
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	DATE	(State in full, without abbreviation) RANK, BRANCH of SERVICE, and ORGANIZATION, specifying Co., Reg., and Div., or similar designation of unit	LOCATION	Abbreviations for RANK BRANCH
Enlistment				
	1 1917	Mar Savings Committee of Illinois	Chicago	auditor
	Feb., 2 1918	Office of the S. M. G. of the army	Mashingto	Head methods Control
	3 1918	1 Ditto	Washington	Chief admr. Div.
Promotions, Transfers,	4 1918	asst. Director of Truance, V.S.a.		
Offices held, e.g., Co. Comdr., Adj., etc.	8 1919	member, Committee on Princis	Madmin	instration
	6	V. S. S. B., Emergency	lest Con	paration
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Discharge				

As Asst. Director of Finance was in command of 361 officers on nov. 11, 1918.

Date Jam. 22, 1.920					
			RANK, BRANCH	DATE	
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	and the	M. E. orthice			
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Name in full	Frank M. Firman	Date
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Academic title	Instructor	The	School	01	Education.

	DA	TE	(State in full, without abbreviation) RANK, BRANCH of SERVICE, and ORGANIZATION, specifying Co., Reg., and Div., or similar designation of unit	LOCATION	Abbi RANK	reviations for BRANCH
Enlistment	Aug	. 1917	2nd Officers Training Camp	Fort Sheridan	Officer Candidate	F.A.
	I Nov	1917	Officers Reserve Corps U.S.A.	Ft. Sheridan	LT.	F.A. OR.C.
	2 Dec		Sailed for Overseas Service	France		
	3 April	1 1918	Field Artillery School, SAUMUR, France	France		
Promotions, Transfers,	4 Luly	-1918	146th Regiment U.S. Field Artillery	Germany United States.		U.S. Field Art.
Offices held, e.g., Co. Comdr., Adj., etc.	5 00	1918	Recommended for promotion at end of MARINE. Campaign August 1918; Promotion October 1918	Nantillois France	Lieutenant	U.S.F.a.
	6 May	1919	Recommended for promotion october 1918; Promotion May 1919	Bendorf Germany	Capt.	U.S.F. a.
	Oct.	1918	146 FA	France Germany	15t Lt.	F.a.
	Mai	rch 1919		Germany U. S.	1st Lt Capt.	F.a.
Discharge	152	luly 19.	Camp Grant, Rockford, 111.		Capt	F. a.

Include all service civil and military

With Regiment in following campaigns:

Champaign-Marne Defensive 15-18 July 1918

Marne-Aisne Offensive 18 July - 12 August 1918

St. Miniel Operations 12 Sept - 15 Sept. 1918

Argonne-Meuse Offensive 26 Sept - 11 Nov. 1918

American Army of Occupation - Germany December 1918 to

	DATE	(State in full, without abbreviation) RANK, BRANCH of SERVICE, and ORGANIZATION, specifying Co., Reg., and Div., or similar designation of unit		
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		Regimental Contitions Officer		
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Include all service civil and military

December 1918 h

of

Henry Gordon Gale.

Entered Second Officers' Training Camp Ft. Sheridan, August
27, 1917. Commissioned Capt. Infantry, November 27, 1917. Reborted at Camp Grant, December 15, 1917. Detailed as Senior
Instructor, Third Officers' Training Camp, December 15, 1917.
Commissioned Capt. Signal Corps, January 15, 1918 and reported
at Washington, D.C. Commissioned Major Signal Corps and sailed
for Europe, January 31, 1918. Arrived at Liverpool, February
16th. At LeHarve, February 19th. Stationed at Paris and Langres
until June 15th. Stationed at Tours as officer in charge of the
Meteorological Division and later Officer in Charge of the Special
Service Division, Office of the Chief Signal Officer. Commissioned
Lieut. Col. Signal Corps, March 17, 1919. Left Tours March 19, 1912
Left Brest, March 26, arrived at Hoboken, April 2, reported at Camp
Alfred Vail, April 3. Reported at Washington, D.C., April 4. Honorably Discharged April 4, 1919.

Decoration as Chevalier of the Legion of Honor conferred by the President of France, April 4, 1919.

Citation by the Commander in Chief of the American Expeditionary Forces for "exceptionally meritorious and conspicuous services", June 3, 1919.

20

Henry Gordon Gale.

Entered Second Officers' Training Camp Ft. Sheridan, August 27, 1917. Commissioned Capt. Infantry, November 27, 1917. Reported at Camp Grant, December 15, 1917. Detailed as Senior Instructor, Third Officers' Training Camp, December 15, 1917.

Commissioned Capt. Signal Corps, January 15, 1918 and reported at Washington, D.C. Commissioned Major Signal Corps and sailed for Europe, January 31, 1918. Arrived at Liverpool, February 16th. At LeHarve, February 19th. Stationed at Paris and Langren until June 15th. Stationed at Tours as officer in charge of the Service Division and later Officer in Charge of the Special Meteorological Division and later Officer in Charge of the Special Lieut. Col. Signal Corps, March 17, 1919. Act Tours Service Demissioned Left Breat, March 26, arrived at Hoboken, April 2, reported at Camp Orably Disciparged April 4, 1919.

Decoration as Chevalier of the Legion of Honor conferred by the President of France, April 4, 1919.

Citation by the Commander in Chief of the American Expeditionary Forces for "exceptionally meritorious and conspicuous services", June 3, 1919.

Name in fullAcademic title	Idas a. M	us Ravid Gaebler 1.	Date	4/9/20
	DATE	(State in full, without abbreviation) RANK, BRANCH of SERVICE, and ORGANIZATION, specifying Co., Reg., and Div., or similar designation of unit	LOCATION	Abbreviations for RANK BRANCH
Enlistment	3/9/18	Pt 1stales Plato Seat 36 Ox SWOW	Rochesty.	Part Photo #36 6

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3/19/19

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State in detail military exploits performed, honors and marks of distinction conferred (including service chevrons), and any items of interest concerning your service, such as position of instructor in special subject, etc.

Include all service civil and military

information for The University of Chicago Telas Borbus

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	LOCATION RANK BRANCH					
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			DATE RANK, BRANCH of SERVICE, and ORGANIZATION, specify Co., Reg., and Div., or similar designation of unit 3/9/18 Pot. 1st class Photo Sect. 36 Apr. 5/9/18 5/9/18 1/9/18 2 2 4 4 4 4 4 4 4 4			
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· Date 4/9/20

ploits performed, honors and marks of distinction conferred (including service chevrons), and any items of interest concerning your service, such as position of instructor in special subject, etc.

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Name in full_ Academic title_		Cha	sociate Professor of Germanic	Date Jan. 17,			
Academic title_			Topen to by the manne	e margy.			
		DATE	(State in full, without abbreviation) RANK, BRANCH of SERVICE, and ORGANIZATION, specifying Co., Reg., and Div., or similar designation of unit	LOCATION	Abbreviations for RANK BRANCH		
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Offices held, e.g., Co. Comdr., Adj., etc.	5						
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Discharge							

I read the files of German numbers published in Illinois, Fraconain and Indiana, under the direction of the Chicago Poor Office authorities for the purpose of investigating the war records of these papers (on the score of patriotism, service, attende toward the war). The period covered about argear and a half previous to the armistice. I received a sollar a year appointment from the Federal Postal authorities shortly before the armistice for this work. were of the last himinaity pelicago dit. Loan committee.

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					relations for BRANCH

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Date February 1920 Name in full_ Academic title (State in full, without abbreviation) Abbreviations for LOCATION DATE RANK, BRANCH of SERVICE, and ORGANIZATION, specifying RANK BRANCH Co., Reg., and Div., or similar designation of unit Ordnance Dysarlment Enlistment why 17, 1918 Promotions, Transfers, Offices held, e.g., Co. Comdr., Adj., etc. Dec6,1918

State in detail military exploits performed, honors and marks of distinction conferred (including service chevrons), and any items of interest concerning your service, such as position of instructor in special subject, etc.

Discharge

Date February 1920		
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]	DATE	(State in full, without abbreviation) RANK, BRANCH of SERVICE, and ORGANIZATION, specifying Co., Reg., and Div., or similar designation of unit	LOCATION	Abbi	eviations for BRANCH
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John Bal and

On the organization of the Draft. Iwar appointed by four howden, as Chairman of Division 16-Chicings, - the Etoselson district. We sent a regiment of men into the service from our division.

Date San 5. 1972

Jack Goods

nformation for The University of Chicago Calar Server

DATE RANK, BRANCH of SERVICE, and ORGANIZATION, specifying LOCATION RANK BRANCH CO., Reg., and Div., or similar designation of unit	
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(including service chevrous),
and any items of interest concerning your service, such as
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subject, etc.

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Name in full Goodnich, Idelau Rees Academic title		delan Reis	School of Ed x	Date			
		DATE	RANK	(State in full, without abbreviation) BRANCH of SERVICE, and ORGANIZATION, specifying Co., Reg., and Div., or similar designation of unit	LOCATION	Abbre RANK	eviations for BRANCH
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ischarge							

her period dering the war was so very slight 700 uno fficial that it would add little or nothing to the lewernest trees. During the time I was an instructor at university of Chieseso I assisted under his Proctor Smith in the war Camp Community Service at the Sailors of Soldiers Club at 20 1st bashing tow Street I was preparing some work in clothing conservation to be used at his heavises Food Conservation Rooms on Wabash are when the armitice cause of the meaterial was never used yours very touly pillen Goodrich.

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Sept. 9, 1919

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work was limited to reading German and Scandinavian newspapers

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Name in full William & Gray

Academic title Associate Professor & Education

(State in full, without abbreviation)

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Name in full_	James Parker Hall Professor of Law; Dean of Law School.	Date	Jan. 28/20,
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State in detail military exploits performed, honors and marks of distinction conferred (including service chevrons), and any items of interest concerning your service, such as position of instructor in special subject, etc.

Nos. 1 and 4 will not be published, as General Seibert is opposed to the publication of No. 1.

A copy of No. 5 is inclosed.

No. 2 is now in print in the Proceedings of the National Academy of Sciences.

The above is probably not worth publishing, but I send it is much as information,

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	11/27/	17 Commissioned First Lieutenant of Infantry.	e		• •
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	6	as Liaison Officer and Assistant Provost Marshall.			Inf.L.S
	7 8/26/1	Stationed as Instructor with French Military	Faverney	60 00	Inf.
	8 3/29/1	Mission at Faverney, Haute Saone, France. * 9 Leave Officer and Interpreter to Post Comm'd't.g. H. Q.	Chaumont		
Discharge	Honorabl dischard	ed At Camp Grant, Sept. 13, 1919.			

State in detail military exploits performed, honors and marks of distinction conferred (including service chevrons), and any items of interest concerning your service, such as position of instructor in special subject, etc.

*Instructor in English and Lecturer on U.S. History, and after the Armistice, Instructor in German to French officers.

Three gold service chevrons.

Include all service givil and military

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			ed First Lieutenant of Infantry to gray in Army			
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discharged At Camp Grant, Sept. 1219.

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The University of Chicago

Office of the President

51 Forest Avenue, East, Detroit, Michigan, Oct. 3rd. 1919.

My dear Ex-Colonel Gale:-

It seems much more than two years ago that you and I were "Cahdidates" together at Fort Sheridan, renewing our youth with the boys. Both of us have gone through a great variety of experiences since then. You, I am informed, returned from France several months ago. I received my honorable discharge only the thirteenth of last month. Perhaps a brief summary of my varied forms of service abroad may interest you.

After the first eleven weeks spent in Officers' training camps in France, of which I gave you an account when your guest at luncheon in the Hotel de la Poste at Langres. I went to Paris to report to the Chief Liaison Officer. I bore with me orders from G.H.Q. assigning me as Liaison Officer with the First Division. A few days after reporting in Paris, the Chief Liaison Officer took me in his autocar to Ist. Div. Hgrs. then located about fifty miles north of Paris (the division being on its march toward the Front near Cantigny). To my deep regret the commanding officer at Hors. refused to accept any additional liaison officers at that time, preferring to wait until he should have something for them to do, so I was conducted back to Paris forthwith. A few days later orders reached me assigning me to duty as Liaison Officer in the Transportation Department of the French Ministry of War, at Paris. There I remained about two months, my duties consisting of translating from English into French (or vice versa) all sorts of orders, requests or communications. I acted also as interpreter at conferences between U.S. and French officers. In the middle of June, 1918, I was ordered to Lyon to be Liaison Officer at the French Regional Military Headquarters located in that city, I had not been at work in this new position two weeks before I was made A.P.M. at Lyon in addition to my Liaison job; and there was no other A.P.M. to share the responsibilities with me. These two positions kept me hustling, I can assure you. Near the end of August, after a sojourn of ten weeks in Lyon, I received a telegram from G.H.Q. ordering me to report to the Adjutant-General as soon as possible. When my successor had arrived and had been broken in, I betook myself with all speed to Chaumont, believing I was at last to be placed in the Intelligence Section or Service, where I was convinced I

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Detroit Avenue Gant. Detroit Michigan.

We sear Ex-Colonel Gale:

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The University of Chicago

CHICAGO, ILLINOIS

Office of the President

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could be of most value. Imagine my surprise when I received orders assigning me as Instructor in English in a school, conducted by the "French Military Mission attached to the American Army", in which French officers (from Lieut-Colonels down to Second Lieutenants), selected on the basis of miditary merit and experience as well as on account of their knowledge of the English language, were being trained to fill positions as Informing Officers with American Army units. This Center of Instruction was located at Faverney (Haute Saone) on the route between Langres and Epinal. The village contained less than 1500 inhabitants, a typically dirty French rural commune. I leave it to you to imagine how enjoyable my new surroundings were after having lived more or less luxuriously in large cities like Paris and Lyon. However, the country round about Faverney is picturesque and healthful and my work proved both interesting and congenial. I was asked to lecture on United States' History as few of the officers. I was informed, had more thatn the vaguest ideas on that subject and all deisred to know more. Of course I lectures in English (or rather in American) and was expected to converse with the French officers in my native dialect on every possible occasion so as to enable them to learn American idioms and pronunciation. I was one of seven American lieutenants giving instruction in English, although each of us lectured on a different subject. The intimate contact with many very intelligent and personally charming French officers, which my new position made possible, was delightful. The commanding officer of the school, a Major (Commandant) of Chasseurs & pied, treated us Americans most hospitably and took an especial fancy to me, making of me almost a boon-companion. He took me on a week's trip with him to his parental home in Le Puy, whence we drove over to visit Lafayette's chateau situated not far distant. He kept me with him at Faverney after he let go the other six American lieutenants until, in fact, the school was disbanded and he himself was given another assignment. That was the end of March of this year. Immediately after the Armistice I was placed in charge of an advanced course in German conversation and when a request was received for an officer able to speak German of English, it was I who examined the officers available for the position to be filled and passed on their linguistic qualifications. When released from Faverney I expected to be sent

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The University of Chicago

Office of the President

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back to the States. But Fate decreed otherwise. A Lieutenant at G.H.Q. was let go home to his wife and I was assigned to fill the vacancy. Thus I became "Leave Officer" at Post Head-quarters G.H.Q. and "Interpreter" to the Post Commandant, Col. Ralston. I was kept on these jobe until G.H.Q. was closed and abandined by the U.S. Army. A 14 day leave being then due me. I took it. my orders instructing me to report at Brest at its expiration for transportation as a casual officer to the U.S. On August 2nd I reached Brest where I waited in idleness for sometime till finally I was appointed second in command of a Brest Casual company of 151 enlisted men. After several days and nights of work with service records, equipping the men, and seeing them through divers kinds of inspections. we were at last ordered on board the U.S.S. Troy on which we had a slow but rather pleasant homeward voyage. At Camp Merritt we had more work with service records and additional inspection nuissancrs. The last lap was entered upon when the company of 300 men, to which I was attached and which was bound for Camp Grant, was ordered to entrain for that destination.

During my eighteen months service in the A.E.F., I failed to report for duty only 15 days on account of illness, viz. three days at Langres when supposed to have the "pink eye" and twelve days spent at Base Hospital 17 at Dijon with the "Flu". I had also two 7-day and one 14-day leaves.

It is not my present intention to return to the History Department of the University of Chicago. During my sojourn abroad, I conceived the idea that I could possibly be of greater value to my country in the Consular Service, inasmuch as I have now a good command of both the French and German languages, have already spent more than seven years of my life in Europe and am more familiar with European conditions than most candidates for consular positions. I shall have to pass a rather stiff examination in difficult subjects, like International Law, Commercial Law, Marine Law, Political Economy and Commercial Geography, but, before this examination is held, I hope to have acquired a sufficient mastery of these subjects to stand the test.

Now the Department of St te informs me that "letters from competent and responsible persons must be filed with the Department of State attesting to the moral character, integrity, good deportment and ability of the applicant and his qualifications" and that "in order that

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The University of Chicago

Office of the President

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due weight may be given these recommendations, the persons vouching for the ap licant's qualification....should fully state the facts which enable them to do so".

You have known me long and well enough, I am sure, to write such a recommendation for me. I do not wish to put you to any great inconvenience. If, however, you could find time in the near future, and if you would be willing to thus vouch for my character and ability, I would be deeply grateful to you. If you decide to write such a letter of recommendation, please address it to the Department of State, Washington, D.C., but send it to me, at the address given above, so that I can send it along with my application. Naturally I desire to file my application and letters of recommendation as soon as possible.

Thanking you cordially, in advance, for this great favor, and extending to you my kindest regards, I am

Very sincerely yours,

Andrew Edward Harvey.

The University of Chicago
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with 34 & Division, decturer in the Divisional School of Hypiens and

Academic title Professor of Anatomy

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	3	Sept.	Sent to Office Div. Surgeon, 34th Div., Deming, N.M.	Camp		
	4	see	Commissioned Major M.C.		Maj	M.C.
Transfers, Offices held, e.g., Co. Comdr., Adj., etc.	5	1918 Jan	Sent to join Base Hospital 13, Fort McPherson, Ga.			
Auj., etc.	6	June	Arrived with B.H. 13 at Limoges, France			
	7	Dec 29	Left Limoges for U.S.A.			
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Discharge		Feb. 14	at Camp Dix, N.J.	•	Maj.	M.C.

State in detail military exploits performed, honors and marks of distinction conferred (including service chevrons), and any items of interest concerning your service, such as position of instructor in special subject, etc.

with 34th Div. at Camp Cody, acted as:Commandant, Schools Medical Dept.; Instructor, Manual Medical Department; Lecturer, School of Hygiene and \$ Sanitation of the Division; Assistant Sanitary Inspector; Mess Officer, Div. Hq. Mess; Member Disability Board; Consulting Surgeon, Orthopedic Casual Detachment; Member of Board for the preliminary examination of candidates for commission, Med Dept; Medical member Board of Officers for the examination of candidates for commission, Q.M. Corps:

with Base Hospital 13: Instructor, Manumal Medical Dept.; Registrar; Receiving Officer; Summary Court; President Special Court Hospital Centre of Limoges.

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Name in full Hektoen, Ludvig	Date January 16, 1920
Academic title Professor of Pathology	

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Did not hold any military appointment, but acted as consultant in infectious diseases at U. S. Naval Training Station at Great Lakes, Ill., during 1917 and first half of 1918; was director of Red Cross Laboratory Car Metchnikeff until the car was taken over by the army; prepared articles on the bacteriology of measles and on experimental measles at the request of the Surgeon-General's office; helped to formulate standard technic for detection of meningococcus in military camps; supervised examination for sterility of catgut by officers of the sanitary comps U.S. army.

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VARIATIONS, DUE TO HEAT TREATMENT, IN THE RATE OF ADSORPTION OF AIR BY COCOANUT SHELL CHARCOAL¹

By HARVEY B. LEMON

RYERSON PHYSICAL LABORATORY, UNIVERSITY OF CHICAGO Communicated by A. A. Michelson, May 19, 1919

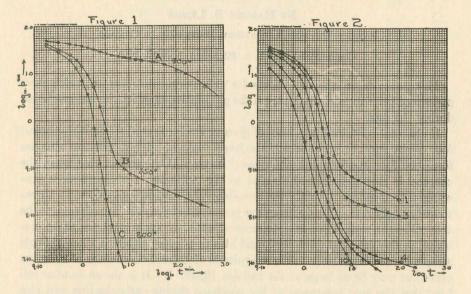
This paper is an abstract of results of experiments most of which were performed prior to January 1918, but the publication of which has been withheld during the war.¹

If charcoal contained in a bulb of glass or iron is heated to 600°C for about four hours and the gases which are freed pumped off into the low vacuum furnished by a mercury diffusion pump,² it is a well known fact that when the charcoal is then subsequently cooled to the temperature of liquid air it possesses a tremendous capacity for adsorbing gas, and will adsorb it at an extremely rapid rate. This adsorption power, however, has been found to be very different with different specimens of charcoal made from the same material, in this case cocoanut shell, and also to be very different when a single sample is used repeatedly. The magnitude of these differences is of no mean order but may be as large as the ratio of 10,000:1. It is shown in what follows that the heat treatment of the specimen during carbonization and also during successive 'outgassings' is a decisive factor in the control of the efficiency of the material as an adsorbent.

The experimental method was one of extreme simplicity. The shell was carbonized in an enclosed electric furnace having a vent for escape of gases and vapors. The temperature was indicated on a Leeds Northrup potential point resistance thermometer. The charcoal was then ground up to particles of from 1 to 3 mm. diameter and cleansed from all smaller fragments and dust. A definite weight (25.7 grams in most of the experiments, determined in dry air with which the charcoal was saturated) was sealed up in a tube of iron, quartz or Pyrex glass depending on the temperatures to be subsequently used. From this tube cocks communicated, (1) to the diffusion pump, and (2) to a fixed volume that could be filled with dry air at any desired pressure. This fixed volume included a McLeod gauge and a mercury barometer column so that the pressure in it could be read to within a few per cent over a range

of from 100 cm. to 0.00001 cm. A small Geissler tube was in communication with the charcoal bulb for a rapid means of observing the character of the gas content and, qualitatively, the pressure. After outgassing through cock, 1, this was closed and cock, 2, opened. The charcoal was thus exposed to a constant volume of 873 cc. initially filled with air at about 90 cm. pressure. This initial pressure was varied slightly as the room temperature varied so that the enclosed mass of air was constant, i.e. 1.2 grams. The air was immediately adsorbed and the rate of fall of the pressure observed as long as it was appreciable.

Results are expressed graphically by plotting the logarithm of the pressure in cm. against the logarithm of the time in minutes elapsing after the cock, 2, was opened. Figure 1 shows curves for three different samples under



identical conditions. A was carbonized at 900°C., B at 850°C. and C at 800°C. They were all outgassed simultaneously for 6 hours at 425°C. and tested in rapid succession on the same apparatus. The enormous difference in rate is obvious in view of the logarithmic scales. The initial pressure of 90 cm. is reduced in 10 minutes to 20 cm., 0.71 cm. and 0.0003 cm. for A, B and C respectively.

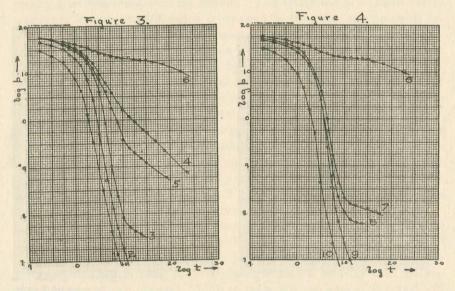
Figure 2 shows the results of successive 'runs' on the same specimen. This specimen was carbonized at 850° C. and the outgassings were all identical at 600° for $4\frac{1}{4}$ hrs. It shows a systematic increase in rate after each run and subsequent outgassing becoming equal after 10 consecutive runs to the former sample shown as C in figure 1.

Figure 3 shows an effect which was controlled so as to be the converse of that illustrated in figure 2. It depicts also successive runs on a single speci-

men, the carbonization of which was for 3 hours at 670°C. The first run shows an extremely active adsorption. Successive outgassings were not now identical but were as indicated below:

First	 4.5 hours at 633°C.
Second	 0.5 hours at 800°C.
Third	 0.5 hours at 850°C.
Fourth	 0.5 hours at 875°C.
Fifth	 4.0 hours at 640°C.
Sixth	 0.7 hours at 905°C.

The first four runs show a cumulative loss of activity as result of high outgassing temperature. In the fifth the activity is in part restored by prolonged outgassing at the lower temperature and in the sixth it is again almost



totally destroyed. Quartz tubes are unsuitable for use for the high temperature work because of their devitrification by the hot carbon vapor.

Figure 4, repeating the sixth run shown in the preceding, carried the same sample back to high activity in four subsequent outgassings which were as follows:

Seventh	22.0 hours at 650°C.
Eighth	22.0 hours at 500°C.
Ninth	44.0 hours at 650°C.
Tenth	1.0 hours at 840°C.

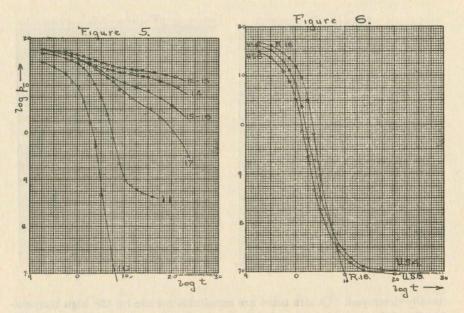
A sluggishness of behavior is apparent after this much use as indicated in the long times necessary to produce a change in quality and also in the failure to respond as before to a temperature higher than 800.

Figure 5 shows continuation of the attempt for the second time to destroy the activity. Repeating curve 10 of the preceding and outgassing

Eleventh	0.5 hours at 950°C.
Twelfth	1.5 hours at 985°C.

the activity is again reduced. Recovery is now as before sluggish and was not carried to completion.

Thirteenth'	 3.0 hours at 650°C.
Fourteenth	 8.0 hours at 625°C.
Fifteenth	 .12.0 hours at 645°C.
Sixteenth	 1.5 hours at 650°C.
Seventeenth	 .14.0 hours at 650°C.



The characteristic form of the family of curves shown is undoubtedly of generic quality. These curves have been reproduced on a wide variety of samples of cocoanut shell charcoal many times; and when the results are assembled exhibit almost every degree of gradation from a central, closely linear (after the first minute) relationship. Similar experiments on pure gases rather than air are in progress in the hope of simplifying this family of curves somewhat so as to obtain an interpretation expressible in analytic form.

Two hypotheses have been advanced for these phenomena. One ascribes the changes in the sense of increased activity to the gradual distillation out of the material of heavy nonvolatile hydrocarbons, many of which in the form of tars and gums come off during the initial carbonization. The other is based on the conception advanced by Miss Ida Homfray³ that these phe-

nomena are to be regarded as saturated solutions of a rigid phase, carbon, in a fluid phase, gas. It suggests that repeated solution and subsequent evaporation, here taking the form of adsorption at low temperatures and outgassing at high, produces a gradual modification of the character of the carbon with respect to its fineness of division, something similar to precipitation occurring.

Loss of activity seems difficult to account for on the first hypothesis. It is to be noted that permanent loss of activity can always be produced on any sample by heating to 1200°C. This is usually ascribed to a partial destruction of porosity and has been observed before. Attempts to extract heavy hydrocarbons by the use of the lightest liquid solvents, ligroin alcohol and acetone were inconclusive. Considerable amounts of tarry material were removed, but not by the solvents directly. They distilled out during outgassing after treatment with the solvent. This treatment in all cases causes temporary loss of activity which is renewed in the usual manner with repeated use after low temperature outgassings. Experiments by others however⁴ which have shown that activation is possible by other methods than the ones here outlined seems to favor the hydrocarbon hypothesis.

Those other methods for activation of field material for the adsorption of complex vapors have been used in conjunction with the above described process. A comparison of material activated by use and low temperature outgassings alone as herein described, with the most highly activated charcoals produced elsewhere under the conditions of these experiments is given in figure 6. The crossing of the curves may be very significant. The laboratory charcoal, R. 16, is less active initially but ultimately runs to lower values than the field material U. S. 4 and U. S. 6. The differences at either end are not large in comparison with the range of the phenomena discussed above.

Experiments are in progress on saturation values for adsorptions of mixtures of varied proportions. An hypothesis originally advanced by McBain⁵ that there is a distinction to be made between surface condensation and interior diffusion is also being subjected to experimental scrutiny with modern materials now at our disposal. Both of these lines of work it is hoped will shed light on the mechanism of the process.

A more detailed account of this work will shortly appear in the Physical Review as a series of papers under the general title of Studies in Charcoal Adsorption.

¹ This article is published with the approval of Major General William L. Siebert, Director Chemical Warfare Service, U. S. A.

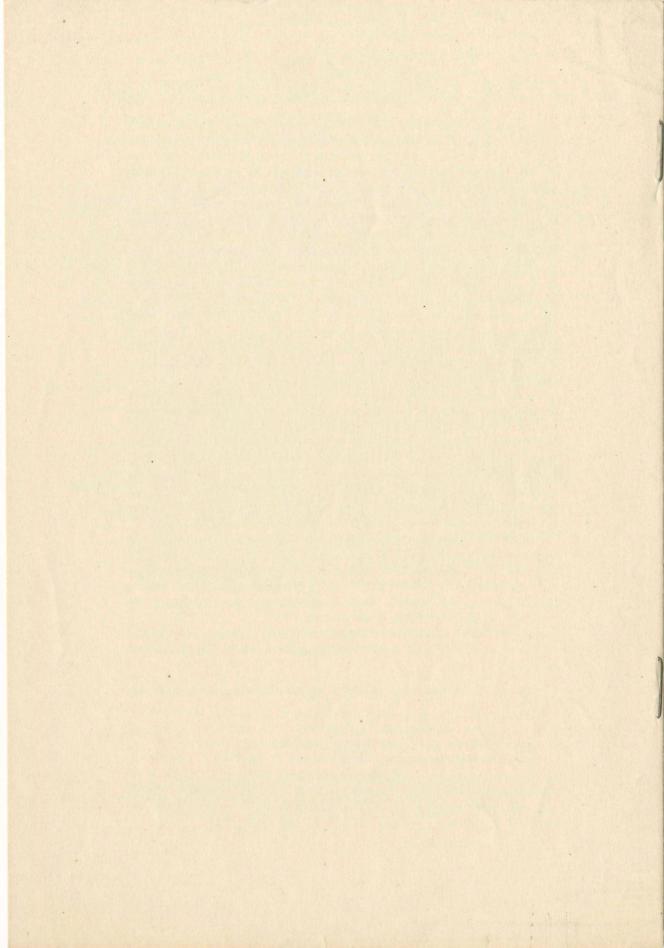
² Shrader, Ithaca, N. Y. Physic. Rev., 12, 1918, (70).

³ Homfray, Liepzig, Zs. Phys. Chem., 74, 1910, (139).

⁴ Dorsey, Easton, Pa., J. Ind. Eng. Chem., 11, 1919, (284)

⁶ McBain, London, Phil. Mag., 18, 1909, (916).

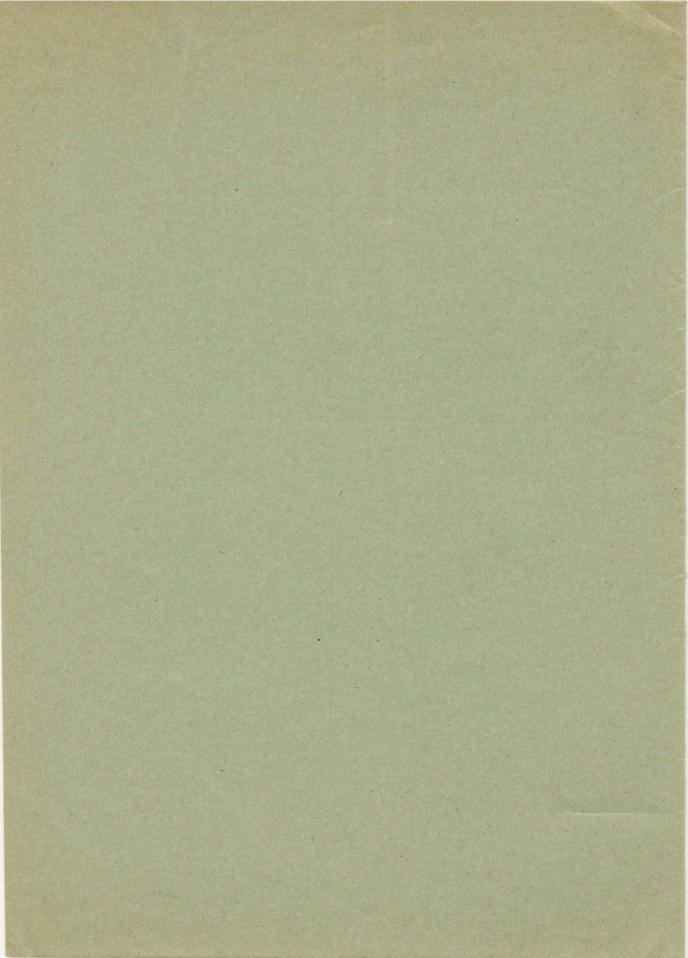
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STUDIES OF THE ADSORPTION OF GASES BY CHARCOAL. I.

By HARVEY BRACE LEMON.

[Reprinted from the Physical Review, N.S., Vol XIV, No. 4, October, 1919.]



STUDIES OF THE ADSORPTION OF GASES BY CHARCOAL. I.

BY HARVEY BRACE LEMON.

VARIATIONS DUE TO HEAT TREATMENT, PRE-EQUILIBRIUM EFFECTS.¹

SYNOPSIS.

Problem.—By varying the heat treatment of a given specimen of cocoanut shell charcoal both the rate at which it adsorbs air and the total amount it can adsorb from a given small mass of air can be varied over a wide range. These variations are studied for the case in which the mass of air used is less than that required to saturate the charcoal.

Activation.—An increase in the adsorptive power, "activation," is produced by repeated evacuations at 650° C., each evacuation being followed by an adsorption of air at the temperature of liquid air.

De-activation.—A decrease in the adsorptive power can be produced by increasing the evacuation temperature to between 800° and 900° C.

Hypotheses.—Hypotheses advanced to explain these phenomena are discussed.

EARLY in 1915, while making up some demonstration apparatus for the purpose of showing the presence of the noble gases in the atmosphere using the method of Gehlhoff, large differences were observed in the efficiency of the carbon employed on different occasions. Reference to the literature showed a very great bibliography on the subject of carbon with especial reference to its use for adsorption of gases but no light was shed on the particular question at issue which was why different samples of charcoal made from the same material could show variations in the ratio of 1:10,000 in the amount of gas that they would adsorb in a given time. Experiments were begun therefore which resulted in two reports on the subject, the latter of which was presented at the

¹ This article is published with the approval of Major General William L. Sibert, Director Chemical Warfare Service, U. S. A.

² Phys. Zeit., 14, 838, 1913.

Chicago meeting of the American Physical Society, December 2, 1916, and published in the Proceedings in April. In May, 1917, at the request of Mr. Bradley Dewey, then assistant in charge of the Investigation of Gases in Warfare, further publicity of this work was indefinitely postponed. Experiments have been continued on the subject ever since as time has permitted. The departure of every one of the senior members of the staff of this laboratory into the service of the Navy or the Army left the writer with but little time for research and later a call for his services in the Ordnance Department took him away from it altogether. Many others have in the meantime become interested in this subject and a great amount of valuable information has been obtained, especially by some whose previous experience made them especially fitted for work of this sort.

The present paper deals with the early experiments on the subject most of them prior to January, 1918.

The method of experimentation adopted was one which made possible the determination of the time rate of adsorption under the conditions of constant volume. From a practical point of view this seemed desirable. Experimentation with this method falls into two broad categories. (a) By using a relatively large amount of gas, conditions of saturation can easily be obtained, and the time necessary to produce them observed. (b) By using a relatively smaller amount of gas so that saturation may never be reached very striking differences in the rate of adsorption of different samples or of the same sample under different conditions become apparent, and it is hoped that this may throw further light on the nature of the process. This latter method of use is herein described under the title of "pre-equilibrium effects."

The apparatus is simple and the method of experimentation may be described with reference to Fig. 1. A McLeod Gauge, M, having three different ratios (18,920–806.1–107.5) connected to a mercury manometer, H, formed a convenient method of observing pressures over the range covered, *i.e.*, 0.00001 cm. to 100 cm. The charcoal was contained in tubes, C, either of iron, quartz or Pyrex glass, depending on the temperature to which it was desired to heat it. Small Geissler tubes were attached to the charcoal bulbs for the convenience of identifying the gas content and affording a quick qualitative indication of the pressure. A volume, B, closed from the tubes by carefully selected stopcocks, served as a reservoir to hold the gas whose rate and total adsorption was measured by the gauges M and H. The system was evacuated either by Gaede pump (drum rotating in mercury) or by tandem mercury

¹ PHYS. REV., 9, 336, 1917.

aspirator and diffusion pump, backed by water aspirator.¹ The three charcoal tubes could be heated and evacuated, "outgassed," simultaneously by an electric furnace. Temperatures were measured by a Leeds & Northrup potential point resistance thermometer with bridge calibrated in degrees C.

The experimental method was to outgas the charcoal for measured

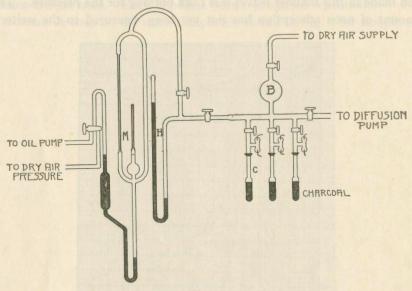


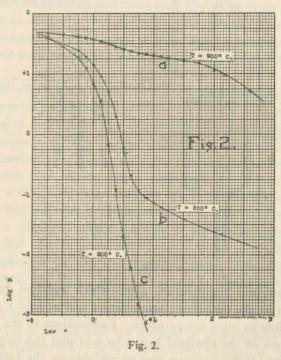
Fig. 1.

time and measured temperature after which residual pressure was observed both at the outgassing temperature and at 500° C., at 400° C. and after cooling to room temperature. Connections to the pumps were shut off as soon as the charcoal began to cool and the outgassing considered complete if the pressure was unreadable on the highest ratio of the gauge after cooling to room temperature. The bulb B and the gauge space was then filled with dry air at a pressure of about 90 cm. (This varied somewhat with the room temperature in order that the mass of gas should be constant = 1.2 gms.) The mass of charcoal used in each tube was 25.7 gms. This weight was taken with the material saturated with dry air at 29° C. and 75 cm. pressure.

After being outgassed one of the three charcoal tubes was immersed in liquid air for about half an hour to assure temperature equilibrium, then the stop cock connecting to B was opened and the pressures read

¹ The mercury pumps were obtained through the courtesy of Mr. C. E. Skinner, Westinghouse Electric and Manufacturing Co., East Pittsburgh. They have been described in the literature by Shrader, Phys. Rev., 12, 70, 1918.

at intervals while the contents of B was being adsorbed. Corrections were made for the partial pressure of the residual neon when this began to introduce an error, i.e., at pressures below 0.001 cm. It was assumed that the neon was not adsorbed and a constant amount subtracted. This assumption is obviously not true since pressures frequently are observed which are very much less than the neon partial pressure so that the correction made in this manner leaves less than nothing for the pressure. The amount of neon adsorption has not yet been measured to the writer's



knowledge and pressures lower than 0.0005 are not regarded as at all accurate and not plotted. The interesting phenomena can be produced at very much larger pressures than this.

For the sake of condensation the results are here given graphically. The logarithm of the pressure in centimeters is plotted against the logarithm of the time, measured from the moment of opening the cock, in minutes. This results always in the same characteristic family of curves of which Fig. 2 is a typical example.

The relative amount of gas is not sufficient in any case to produce saturation of the charcoal and the pressure falls throughout the course of the observations, *i.e.*, for over five hours in the slower samples.

Curves a, b, c are three different specimens of charcoal differing only

in their carbonization temperatures which were 900°, 850° and 800° respectively. The outgassing took place at 425° and was continued about six hours. The enormous difference in effectiveness of these three samples is striking. The pressure initially being about 90 cm. is reduced in 10 minutes to 20 cm. in the case of a, to 0.071 cm. in the case of b, and to about 0.003 cm. in the case of c. These curves are by no means unique ones. Each type has been reproduced repeatedly.

Successive runs on the same sample are shown in Fig. 3 for sample 12.

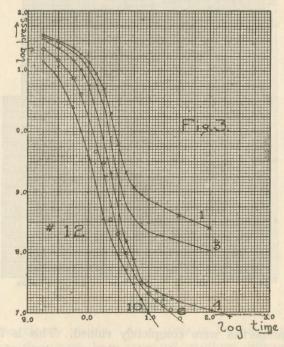
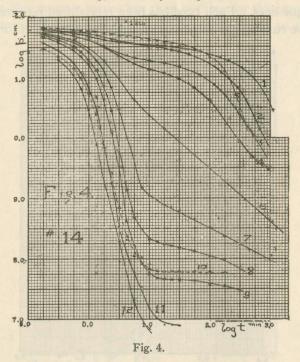


Fig. 3.

This material was carbonized at 850° for 1 hour. Outgassings were at 600° and were never less than $4\frac{1}{2}$ hours which was later adopted as entirely sufficient for complete evacuation. This means only such evacuation as will leave no readable pressure on the highest ratio of the gauge after cooling to room temperature. Residual pressures after such outgassing were at 600° C., 8×10^{-8} cm.; at 400° C., 8×10^{-6} cm. Systematic improvement, *i.e.*, increase of rate, in the sample after ten consecutive runs occurs and after this much use the sample is practically as good as that one shown in Fig. 2, curve c, which was carbonized at a lower temperature. Fig. 4, sample 14, shows a still more extreme case of this variation in a single sample with use. Carbonization was here

made at 875° for $1\frac{1}{2}$ hrs. outgassings, at about 600° for $4\frac{1}{4}$ hrs., as in Fig. 3. The entire range of behavior from that of a very poor sample to that of a very good sample is covered in 12 consecutive runs. Curve 5 is irregular and due to a leaking stop-cock as is also the latter part of curve 10.

It was observed if at any time any sample was heated to 1200° its



adsorptive properties were completely ruined. This is illustrated in Fig. 4, A, which is a continuation of work on sample 14 of Fig. 4. The volume of gas admitted was here increased from 873 c.c. used in all previous curves to 1,707 c.c. and the first curve 13 shows saturation obtained in about 34 minutes. The sample was now put in a transparent quartz tube. Several outgassings were made at 600° and slight improvement resulted, the saturation value falling considerably, the seventeenth run being the lowest. Between the nineteenth and twentieth runs the temperature was raised to 1200° for 20 minutes. The twentieth run is at the top of the sheet showing the almost complete loss of adsorptive power. On reducing the volume of gas admitted to 351 c.c. and taking several more runs the curve remained stationary in this location.

The transparent quartz tube was attacked by the carbon at this high temperature and completely altered in its physical characteristics.

It was almost completely devitrified and became of about the appearance of egg shell opaque and brownish in color with little mechanical strength. It was left with sufficient strength however to withstand pressure and was not observed to leak. The transformation is rather peculiar in not becoming visible until the tube cools. On being taken out of the furnace it is a bright cherry red and perfectly transparent, the charcoal in the

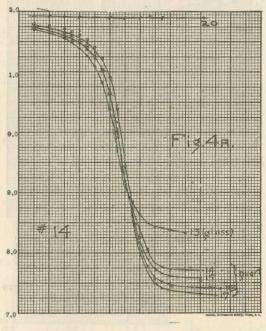
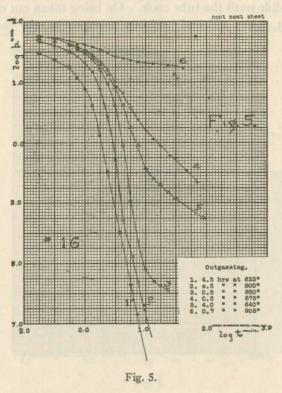


Fig. 4A.

interior being clearly visible. As it cools in the air it grows opaque. These observations have been repeated on three different tubes with identical results. The change is ascribed to a reaction between the carbon and the quartz although this reaction is commonly supposed to occur only at temperatures considerably higher. The temperature is much too low to allow of the formation of carborundum and it is thought that this compound may be identical with a substance known as silfrax which has been described as having similar physical properties to those here observed and which is found in the cooler portions of the carborundum furnaces. Because of the failure of quartz in this respect iron tubes were adopted in future work at higher temperatures.

It was next found possible to change the adsorptive power of a sample in the reverse of the manner shown above. Fig. 5 of sample 16 is illustrative. This sample was carbonized three hours at 670°, just a trifle

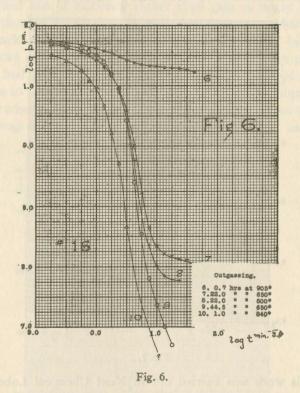
above the temperatures to be used subsequently in normal outgassing. The first run indicated an admirable efficiency. The method of reducing the efficiency was to increase the outgassing temperature to 800° or higher and the details of outgassing between each consecutive run are



given in the insert on the figure. The curves rise for four runs with increasing the outgassing temperature each time. By reducing it to 640° and prolonging the time the sample shows marked recovery. A short 900° outgassing pushes the curve to the top of the page. Fig. 6 is a continuation of the same set of experiments. Here this same sample no. 16 is entirely restored by a series of prolonged low temperature outgassings. The modern term for this process is activation. Fig. 7 shows further continuation of work with the same material in which it is again rendered inefficient (curves for runs 10 to 13) and finally reactivated again (15 to 17). Changes become less easy to produce towards the end of a long series of runs of this sort. Heating samples in iron tubes to 1200° C. was found to destroy the adsorptive power beyond recovery just as well as when the containing tube was quartz.

The most effective charcoal produced up to the present time for the

adsorption of air under the conditions outlined above has been the material carbonized at 670° and then put through a series of "runs" and "outgassings" which consist of alternately adsorbing 0.00467 gms. of air per gm. charcoal at -188° C. and removing it by heating for $4\frac{1}{4}$ hrs. at 600° C. Two hypotheses have been advanced for the interpretation of this fact. The first is that after carbonization there remain



considerable amounts of heavy hydrocarbons "clogging the pores" of the material and that these are gradually removed by repeated flushing out with gas. The second has been suggested by Sgt. H. C. Thompson, after reflection on the point of view advanced by Miss Ida Homfray, who has done under the direction of Professor Ramsay some of the best quantitative work yet published on the pressure-absorbed volume-temperature relations of a particular sample. From similarity of the empirical relations discovered to those known to hold for concentrated solutions, the process is visualized as if it were a saturated solution of carbon in essentially liquefied gas. Repeated solution and subsequent evaporation (here taking the form of adsorption at low temperature and

¹ Zeit. Phys. Chem., 74, 139, 1910.

outgassing at moderately high temperature) may result in the gradual alteration of the character of the carbon with respect to its fineness of division, something similar to precipitation taking place. As a test of the hydrocarbon hypothesis attempts were made to improve the behavior of the material by treating it with the lighter liquid hydrocarbons, ligroin, alcohol, and acetone. Through the courtesy of Professor Julius

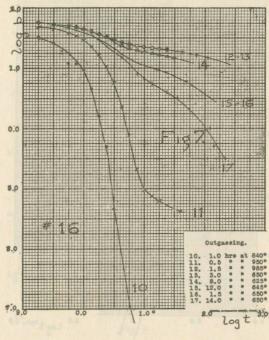


Fig. 7.

Stieglitz this work was carried out in Kent Chemical Laboratory by Mr. M. C. E. Hanke. The liquids were carefully distilled and redistilled to ensure purity. 50 gms. of the charcoal was then boiled in 500 c.c. of the solvent in a reflux condensor in a water-bath for four to five hours when the solvent was filtered off, redistilled and the residues examined. The following residues were obtained:

Samble 21

Ligroin 0.005 gm. Alcohol 0.212 gm.

Acetone 0.38 gm. $\begin{cases} 0.35 \text{ soluble in acetone.} \\ 0.03 \text{ soluble in water.} \end{cases}$

Sample 19.

Ligroin 0.004 gm. Alcohol 0.294 gm. 0.570 soluble in acetone. 0.040 soluble water. 0.010 insoluble in both.

30 gms. of sample 19 extracted with water gave a residue of 0.4 insoluble in a small amount of water.

These charcoals were then subjected to adsorption tests and both showed marked loss of efficiency rather than gain as might have been expected if the heavy clogging materials had been somewhat removed.

Earlier work in which charcoal was extracted with hot hydrochloric acid to remove ash showed similar injury to the adsorptive power. In this case it had been found necessary to wash the material for nearly eight months to get rid of all traces of acid. In all cases of liquid treated charcoal the adsorptive power rapidly recovers with use when the successive outgassings are kept at not higher than about 600° C. as indicated

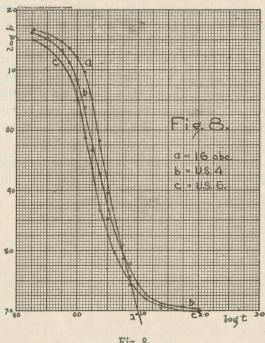


Fig. 8.

in general above. These experiments therefore must yet be regarded as inconclusive not only because of the relative small amounts of impurity extracted but because the subsequent loss of efficiency of the material can readily be interpreted as a secondary result due to the difficulty of freeing the pores from the liquids used.

Other methods of activation based on the hydrocarbon hypothesis

and using gaseous reagents have been developed elsewhere with notable success, especially for field conditions in which the adsorption of a complex vapor is required and where the adsorbing material is already saturated with air, water, vapor, etc. For the rapid production of high vacua, charcoal activated thus is *not* notably better than that discussed above.

A comparison is given in Fig. 8 of the best material produced here by the method of continuous use and that which has been pronounced the best for field conditions in adsorbing noxious vapors and which has been produced by these other methods of activation. The laboratory charcoal, 16 abc, is less active initially but runs to a lower ultimate pressure than does the field material US4 and US6. It is thought likely that there may be a purely surface condensation as well as a slower diffusion into the interior or to less readily accessible surface in accordance with a point of view originally advanced by McBain. Experiments to determine the correctness of this view with respect to modern charcoals are in progress.

No rational expression for the variations in the adsorption described above has yet appeared. In fact even an empirical analytic expression has not been formulated. The phenomena here are of such complexity by nature that it would seem to be more profitable to simplify experimental conditions and secure much more data before attempting to formulate a theory.

Experiments involving the saturation conditions and also experiments on relative adsorption of mixtures are important and will be discussed in subsequent papers.

Ryerson Physical Laboratory, March, 1919.

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